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# **Chapter 11 Review Molecular Composition Of Gases Mixed Answers**

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Composition of Gases ... Chapter 11 - Molecular  
Composition of Gases 11-1 Volume-Mass Relationships  
of Gases I. Measuring and Comparing the Volumes of  
Reacting Gases A. Observations of Gay-Lussac 1. 2  
liters H<sub>2</sub> + 1 liter O<sub>2</sub> → 2 liters H<sub>2</sub>O vapor 2. 2  
volumes H<sub>2</sub> + 1 volume O<sub>2</sub> → 2 volumes H<sub>2</sub>O vapor  
3. 1 volume H<sub>2</sub> + 1 volume Cl<sub>2</sub> → 2 volumes HCl

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4. Chapter 11 - Molecular Composition of Gases 334  
CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol  $H_2$  at STP = 22.4 L  
Oxygen molecule 1 mol  $O_2$  at STP = 22.4 L  
Carbon dioxide molecule 1 mol  $CO_2$  at STP = 22.4 L  
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